**Limiting Reagent / Experimental Error Homework**

Let’s do it all again with this equation:

\_\_\_\_\_ Na2O + \_\_\_\_\_ Pb(NO3)4 → \_\_\_\_\_ PbO2 + \_\_\_\_\_ NaNO3

1) What type of reaction is this?

2) How many grams of lead (IV) oxide will be formed if 0.75 grams of sodium oxide react with 0.80 grams of lead (IV) nitrate?

2) What was the limiting reagent for the scenario in problem 2 above?

3) How much of the excess reagent was left over?

4) If I actually made 0.22 grams of lead (IV) oxide, what was my percent yield? Is it reasonable? Explain why or why not.

5) Did I make any errors when doing this lab? If so, explain what they may have been.